

HW - Problems Ch. 14

14-1 a) P.E. of Falling water \rightarrow Internal Energy
 $\Delta P.E. = mgh = (1.4 \text{ kg})(9.8 \frac{\text{m}}{\text{s}^2})(2.5 \text{ m}) = \boxed{3.4 \text{ J}}$

b) This Internal Energy will appear as increased I.E. of water \Rightarrow higher T

14-9) $1 \text{ W} \cdot \text{h} = \frac{\text{J}}{\text{s}}$, $1 \text{ J} = 1 \text{ W} \cdot \text{s}$

$1 \text{ kJ} = 10^3 \text{ W} \cdot \text{sec} = 1 \text{ kW} \cdot \text{sec} = 1 \text{ kW} \cdot \text{sec} \cdot \frac{1 \text{ hr}}{3600 \text{ s}} = \frac{1}{3600} = 2.8 \times 10^{-4} \text{ kWh}$

14-12) $Q = cm \Delta T$ $\Delta T = \frac{Q}{cm}$

$\Delta T = \frac{125.6 \times 10^3 \text{ J}}{4.2 \frac{\text{kJ}}{\text{kg} \cdot \text{K}} \cdot 5 \times 10^2 \text{ gm}} \quad \frac{\text{kJ}}{\text{kg}} = \frac{\text{J}}{\text{g}}$

$\Delta T = \frac{125.6 \times 10^3}{21 \times 10^2} = (6 \times 10)^\circ \text{K}$ $T_F = T_i + \Delta T = 82^\circ \text{C}$

14-19) $Q = (Cm)_{\text{bottle}} \Delta T + (Cm)_{\text{water}} \Delta T$ $\Delta T = 100 - 15 = 85^\circ \text{C} = 85^\circ \text{K}$

$Q = \left(0.9 \frac{\text{kJ}}{\text{kg} \cdot ^\circ \text{C}}\right) \cdot 0.4 \text{ kg} \cdot 85^\circ \text{C} + \left(4.2 \frac{\text{kJ}}{\text{kg} \cdot ^\circ \text{C}}\right) \cdot 2 \text{ kg} \cdot 85^\circ \text{C} = 30.6 \text{ kJ} + 714 \text{ kJ} = \boxed{745 \text{ kJ}}$

(742 if I used $C = 4.186$)

14-20) $Q = (Cm)_{\text{human tissue}} \Delta T$ $C_{\text{human tissue}} = 3.5 \frac{\text{kJ}}{\text{kg} \cdot ^\circ \text{C}}$
 $\Delta T = 1.4^\circ \text{C}$

$Q = (3.5) \cdot (50) \cdot (1.4) = \boxed{245 \text{ kJ}}$

14-25) H_2 is diatomic $C_v = 20.8 \frac{\text{J}}{\text{mol} \cdot \text{K}}$ Must Find number of moles

We need to convert, P, T, V to standard units

~~P = 10 atm~~ $P = 10 \text{ atm} = 10.1 \times 10^5 \text{ Pa}$, $T = 273^\circ \text{K}$, $V = 250 \times 10^{-3} \text{ m}^3$

$n = \frac{PV}{RT} = 1.1 \times 10^2$ $Q = (C_v)(n) \Delta T = \boxed{57 \text{ kJ}}$

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14-47) a) $R = \frac{d}{kA}$

we want to compare thickness of cork & equiv. air

$d = RkA$

set up ratio between cork & air

$$\frac{d_{\text{cork}}}{d_{\text{air}}} = \frac{Rk_{\text{cork}}A}{Rk_{\text{air}}A}$$

$$\frac{d_{\text{cork}}}{d_{\text{air}}} = \frac{k_{\text{cork}}}{k_{\text{air}}} = \frac{.046}{.023} = 2$$

$d_{\text{cork}} = 2 \text{ cm}$

b) $\frac{d_{\text{tin}}}{d_{\text{air}}} = \frac{66.8}{.023} = 2,900 \text{ cm} = 29 \text{ m}$

14-52) The rate of heat flow thru cork = heat flow thru wood

$$P_w = P_c \quad k_w A \frac{\Delta T_w}{d} = k_c A \frac{\Delta T_c}{d}$$

$k_w \Delta T_w = k_c \Delta T_c$

a) cork inside $\Delta T_c = T - 20$ $\Delta T_w = 10 - T$ $-Tk_w = k_c(T - 20)$

$T = \frac{20k_c}{k_c + k_w} = 52^\circ \text{C}$

b) cork outside

$\Delta T_c = T - 0$ $\Delta T_w = 120 - T$

$k_w(120 - T) = k_c(T)$ $T = \frac{20k_w}{k_c + k_w} = 15^\circ$

c) Since total heat flow is same in both cases thermal resistance is same and it doesn't matter if cork is inside or outside

14-59) $\lambda_{\text{max}} T = 2.9 \times 10^{-3} \text{ m K}$, $T = 1650 \text{ K}$

$\lambda_{\text{max}} = 1.76 \mu \text{m}$

14-63) heat lost = heat absorbed + heat produced

heat produced = 90W

heat absorbed = Intensity \times Area = $(7 \times 10^2 \frac{\text{W}}{\text{m}^2}) \times .57 \times (1.8 \text{ m}^2) \times (.42)$
= 300W

heat loss = 90W + 300W = 390W